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CHM191
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1. Thus using starch as indicator, the colour variations of the solution depend on _____ formation of the starch with iodide ion formed in the last stage of the titration.

--->> Complex

Colloid

Hydrolysed

Navy blue

2. The study of thermal changes in chemical and physical processes is known as _____

thermochemistry

--->> Thermodynamics

A and B

B only

3. In the _____, standard solutions of iodine are used to estimate directly the concentrations of some oxidizable species.

--->> Direct method

indirect method

standard method

calibrated method

4. The starch solution forms a _____ complex with the tri-iodide ion during the titration that is rapidly discharge at the end point

Light blue

Blue

--->> Blue black

Brick red

5. The _____ is the amount of heat released or observed for a given amount of reactants or products

--->> heat of reaction

heat of combustion

heat of neutralization

heat of formation

6. For this reason and in order to avoid complications in the end point detection, the starch indicator is added at the_____ colouration of the solution which is near the end point

Yellow

--->> Light yellow

Colourless

Dark yellow

7. The most commonly used external indicator for iodine titration is_____

Glucose solution

hydrogen chloride

Protein solution

--->> starch solution

8. Of all the oxidizable species available for use, iodine titration are usually used with the____

sulphate ions

--->> Thiopsulphate

hydrogensulphate

thioesters

9. The oxidation potentials of Na is_____

11

23

271

--->> 2.71

10. Complete the colour change in iodometry analysis.
_____(initially)â†’

Red-orange

Dark-brown

Brown

--->> Red-brown